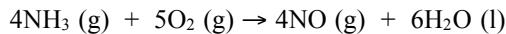


**AP Chem Ch. 5 Study Sheet 5****Multiple Choice***Identify the choice that best completes the statement or answers the question.*

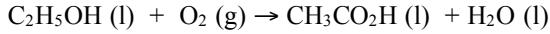
1. Given the data in the table below,  $\Delta H^\circ_{rxn}$  for the reaction



is \_\_\_\_\_ kJ.

Substance	$\Delta H_f^\circ$ (kJ/mol)
$\text{H}_2\text{O}(\text{l})$	-286
$\text{NO}(\text{g})$	90
$\text{NO}_2(\text{g})$	34
$\text{HNO}_3(\text{aq})$	-207
$\text{NH}_3(\text{g})$	-46

- a. -1172
  - b. -150
  - c. -1540
  - d. -1892
  - e. The  $\Delta H_f^\circ$  of  $\text{O}_2(\text{g})$  is needed for the calculation.
2. Given the data in the table below,  $\Delta H^\circ_{rxn}$  for the reaction

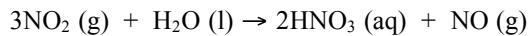


is \_\_\_\_\_ kJ.

Substance	$\Delta H_f^\circ$ (kJ/mol)
$\text{C}_2\text{H}_4(\text{g})$	523
$\text{C}_2\text{H}_5\text{OH}(\text{l})$	-277.7
$\text{CH}_3\text{CO}_2\text{H}(\text{l})$	-484.5
$\text{H}_2\text{O}(\text{l})$	-285.8

- a. -79.0
- b. -1048.0
- c. -476.4
- d. -492.6
- e. The value of  $\Delta H_f^\circ$  of  $\text{O}_2(\text{g})$  is required for the calculation.

3. Given the data in the table below,  $\Delta H^\circ_{rxn}$  for the reaction

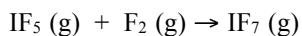


is \_\_\_\_\_ kJ.

Substance	$\Delta H_f^\circ$ (kJ/mol)
$\text{H}_2\text{O}$ (l)	-286
$\text{NO}$ (g)	90
$\text{NO}_2$ (g)	34
$\text{HNO}_3$ (aq)	-207
$\text{NH}_3$ (g)	-46

- a. 64
- b. 140
- c. -140
- d. -508
- e. -64

4. Given the data in the table below,  $\Delta H^\circ_{rxn}$  for the reaction

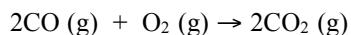


is \_\_\_\_\_ kJ.

Substance	$\Delta H_f^\circ$ (kJ/mol)
$\text{IF}$ (g)	-95
$\text{IF}_5$ (g)	-840
$\text{IF}_7$ (g)	-941

- a. 1801
- b. -1801
- c. 121
- d. -121
- e. -101

5. Given the data in the table below,  $\Delta H^\circ$  for the reaction



is \_\_\_\_\_ kJ.

Substance	$\Delta H_f^\circ$ (kJ/mol)
$\text{CO}$ (g)	-110.5
$\text{CO}_2$ (g)	-393.7
$\text{CaCO}_3$ (s)	-1207.0

- a. -566.4
- b. -283.2
- c. 283.2
- d. -677.0
- e. The  $\Delta H_f^\circ$  of  $\text{O}_2$  (g) is needed for the calculation.